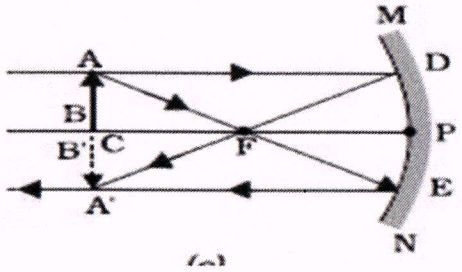
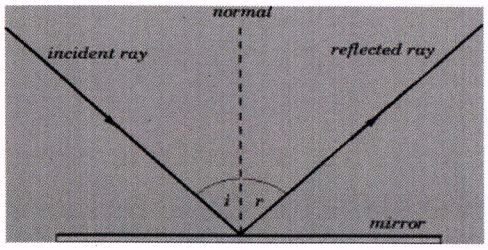
INDIAN SCHOOL MUSCAT

FIRST PERIODIC ASSESSMENT

CLASS X – SCIENCE

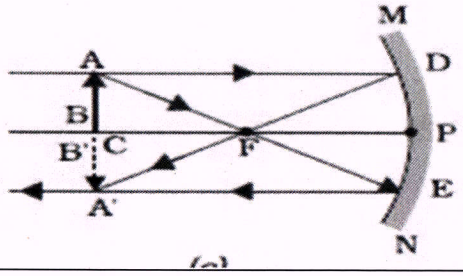
	MARKING SCHEME	
	PHYSICS SET -A	
1.	(a) Concave Mirror (b) Convex Mirror	1
2.	1. A ***real image*** can be caught on a screen whereas a ***virtual image*** cannot be caught on a screen. 2. A ***real image*** is always inverted whereas a ***virtual image*** is always erect. 3. A ***real image*** is formed when the rays of light after reflection actually meet at some point whereas a ***virtual image*** is formed when the rays of light after reflection appear to meet at a point. 4. A ***real image*** is formed in front of the mirror whereas a ***virtual image*** is formed behind the mirror.	2
3.	In case of concave mirror U = -10cm, v = ? m = -3 (real image) m = -v/u v = -30cm	1 + 1
4.		2

PHYSICS - SET B

1.	(i) Convex mirror always gives an erect and diminished image. (ii) Wider field of view.	1
2.	Laws of reflection: Ist law: The incident ray, the reflected ray and normal to the reflecting surface at the point of incidence all lie in the same plane. IInd law: The angle of incidence is equal to the angle of reflection. 	2
3.	V =+ 6.67cm m = +0.56	1 + 1

4.	Ray diagram when an object is placed between pole and focus of concave mirror.	2

PHYSICS - SET C

PHYSICS		
1.	m is positive indicates that the image is virtual and erect. m = 1 indicates that the image is of the same size as of the object.	1
2.		2
3.	1. A real image can be caught on a screen whereas a virtual image cannot be caught on a screen. 2. A real image is always inverted whereas a virtual image is always erect. 3. A real image is formed when the rays of light after reflection actually meet at some point whereas a virtual image is formed when the rays of light after reflection appear to meet at a point. 4. A real image is formed in front of the mirror whereas a virtual image is formed behind the mirror.	4 X ½ = 2
4.	u = -25cm, f = -15cm, v = ? 1/f = 1/v + 1/u V = -37.5cm	1 + 1

MARKING SCHEME OF PERIODIC ASSESSMENT -1 CLASS X (CHEMISTRY)	
CHEMISTRY SET - A	
To justify the law of conservation of mass.	1
In combination reactions, two or more substances chemically combine together to form a single product whereas in decomposition reaction a single substance splits up to give two or more products.	1 + 1
i) Lead Nitrate , Nitrogen dioxide ii) $2Pb(NO_3)_2 \rightarrow 2PbO + 4NO_2 + O_2$	½ + ½ 1
Balance the following chemical equations and identify the type of reaction in each case. a) $2KClO_3 \rightarrow 2KCl + 3O_2$ (Decomposition/Thermal Decomposition) b) $2Al + 3CuCl_2 \rightarrow 2AlCl_3 + 3Cu$ (Displacement Reaction)	½ x 4 = 2

CHEMISTRY SET- B	
Brown coating of copper on iron nail , Blue color of CuSO4 changes to green.	½ + ½
In Displacement reactions, more reactive metal displaces the less reactive metal from its salt solution whereas in double displacement reactions ,exchange of ions takes place.	1 + 1
a) Cathode – Hydrogen , Anode – Oxygen b) H : O in Water is 2 : 1	½ + ½ 1
Balance the following chemical equations and identify the type of reaction in each case. a) $2AgBr$ (sunlight) $\rightarrow 2Ag + Br_2$ (Decomposition/Photolytic Decomposition) b) $BaCl_2 + Na_2SO_4 \rightarrow BaSO_4 + 2NaCl$ (Double Displacement Reaction)	½ x 4 = 2

CHEMISTRY SET - C		
	Brown coating of copper on iron nail , Blue color of $CuSO_4$ changes to green.	½ + ½
	a) Any two observations b) $2FeSO_4 \rightarrow Fe_2O_3 + SO_2 + SO_3$	½ + ½ 1
	In combination reactions, two or more substances chemically combine together to form a single product whereas in decomposition reaction a single substance splits up to give two or more products.	1 1
	Balance the following chemical equations and identify the type of reaction in each case. a) $2KClO_3 \rightarrow 2\,KCl + 3\,O_2$ (Decomposition/Thermal Decomposition) b) $Pb(NO_3)_2 + 2\,HCl \rightarrow PbCl_2 + 2\,HNO_3$ (Double Displacement Reaction)	½ x 4 = 2

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CLASS – X-FIRST PERIODIC ASSESSMENT

MARKING SCHEME – BIOLOGY

SET-A

1.	Carbon di oxide and water.	½+½=1 mark
2.	Villi in large intestine-absorb water from waste products. (½) Villi in small intestine-absorbs nutrients from the digested food. (½)	½+½=1 mark
3.	Present in the leaves (½) Transpiration, exchange of gases (any one) (½) Opening and closing of the stomatal pore is regulated by the entry and exit of water to and from the guard cells (½ + ½ =1 mark)	½ + 1+ ½ = 2 marks.
4.	Breaking down larger fat globules to smaller fat globules by bile. (1) Maintain acidic medium for pepsin to act, kills microbes present in food (½ + ½ =1)	1+1= 2 marks.

SET- B

1.	Chewing helps food to get mixed with saliva and break down it into simpler particles	½+½=1 mark
2.	Carbohydrates (½), starch. (½)	½+½=1 mark
3.	Centre of the guard cells in the epidermal cells of leaves. (½) Transpiration, exchange of gases (any one) (½). When they don't want carbon dioxide for photosynthesis / To reduce transpiration during day time. (any one -1mark)	½ + ½+ 1 = 2 marks.
4.	Contraction and relaxation movement of muscles of oesophagus to regulate the movement of food. (1) Trypsin and lipase (½+½=1)	1+1= 2 marks.

SET -C

1.	Respiration-throughout the day, Carbon dioxide is given out Photosynthesis-only in presence of light, Oxygen is given out. (any one difference)	½+½=1 mark
2.	To break down food into smaller particles (½), to mix food with saliva (½)	½+½=1 mark
3.	Guard cells (½). When it does not need carbon di oxide for photosynthesis and to reduce loss of water (1). Transpiration (½)	½ + 1+ ½ = 2 marks.
4.	Pepsin (½), Digestion of protein (½), Liver secretes bile which helps in emulsification. (1)	½+½+1= 2 marks.